Course Syllabus for CH3510 Fall 2012
Ph Chem I: Thermodynamics, Equilibrium, Kinetics
Department of Chemistry

Instructor Information
Instructor: Loredana Valenzano, PhD, Assistant Professor
Office Location: Room 701A – ChemSci
Telephone: (906) 487-1602
E-mail: lvalenza@mtu.edu
Office Hours: MW 2:30 pm – 3:30 pm, F 2:00 pm – 3:00 pm
Or by appointment

Course Identification
Course Number: CH3510-0A
Course Name: Ph Chem I: Thermodynamics, Equilibrium, Kinetics
Course Location: Room 641 - DOW Envir Sciences & Engr Bldg
Class Times: MWF 1:05 pm - 1:55 pm

Course Description/Overview
To introduce concepts useful in explaining and interpreting the nature of physical and chemical properties of matter. This course will cover the following areas of Physical Chemistry: thermodynamics, chemical equilibrium, and chemical kinetics.

Course Learning Objectives
- To provide students with foundation in thermodynamics principles governing chemical phenomena.
- To guide students in developing quantitative reasoning, problem solving, rigorous thinking but also physical-chemical intuition.

Course Resources
Online Resources
- Canvas: http://courses.mtu.edu
- E-mail List: ch3510-fa12-l@mtu.edu

Required Course Text
Other useful sources may be represented by:

- A comprehensive little wonderful (and cheap) reference is: Enrico Fermi, *Thermodynamics*, Dover Publication Inc., New York, 1936

**Grading Scheme**

Grading System

<table>
<thead>
<tr>
<th>Points</th>
<th>Letter Grades</th>
</tr>
</thead>
<tbody>
<tr>
<td>93-100</td>
<td>A</td>
</tr>
<tr>
<td>87-92</td>
<td>AB</td>
</tr>
<tr>
<td>82-86</td>
<td>B</td>
</tr>
<tr>
<td>76-81</td>
<td>BC</td>
</tr>
<tr>
<td>70-75</td>
<td>C</td>
</tr>
<tr>
<td>65-69</td>
<td>CD</td>
</tr>
<tr>
<td>60-64</td>
<td>D</td>
</tr>
<tr>
<td>0-59</td>
<td>F</td>
</tr>
</tbody>
</table>

*Note that NO CURVING will be applied!*

Grading Policy

Your grade for this course will be based on the following:

<table>
<thead>
<tr>
<th>Assignment</th>
<th>Max points per type of assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>Weekly Online Quizzes (13 – 5 points each)*</td>
<td>50*</td>
</tr>
<tr>
<td>Homework (4 – 25 points each)</td>
<td>100</td>
</tr>
<tr>
<td>Mid-Term Exams (3 – 25 points each)</td>
<td>75</td>
</tr>
<tr>
<td>Final Exam (1 – 50 points)</td>
<td>50</td>
</tr>
<tr>
<td><strong>Total Points</strong></td>
<td><strong>275</strong></td>
</tr>
</tbody>
</table>

* Only your best 10 quizzes will be considered for your final grade.

** Your total score in the Course will be converted to 100%.

Late Assignments

No late assignment will be considered.
Attendance

Attendance in class is not compulsory but it is highly recommended to keep track of the topics covered, and to develop good reasoning skills.

Course Policies

Your grade will be based on:

- **13 online quizzes** assigned via Canvas. At the end of the course, only the best 10 quizzes will be considered for your final grade. This is an attempt to balance missed assignments, silly mistakes, and so forth. Note that the weight of these online quizzes equals almost 20% of your Final Grade in the Course (!) therefore my friendly suggestion to you is to try not to miss them;

- **4 homeworks**;

- **3 midterm exams**;

- **1 final term exam**.

Weekly, I will assign **13 online quizzes via Canvas every Thursday**. Typically, you will be asked to answer 5 questions. You will have a 24 hours window to access the quiz and solve it in 10 minutes, so that you’ll have on average 2 minutes per question as during Exams.

**The 4 Homeworks** will be assigned **roughly every 2 weeks** and will be collected in class the same day one week after (example: homework assigned on Friday, 7th will be due on Friday, 14th). No late homework will be accepted. Students who will not turn in homework without documented and satisfactory explanation will receive a grade of 0.0 (see the last page of this document “Excused Absence”).

**The 3 Midterm and Final Exams** will be assigned in the form of **multiple-choice questions** where you will be asked both to show the development of your reasoning skills and your capabilities in solving and deriving equations to solve specific thermodynamic problems. Students who will not show up for Mid-Term and Final Exams without documented and satisfactory explanation will receive a grade of 0.0 (see the last page of this document “Excused Absence”).

Laptops are allowed in class only, not during exams. The following are not allowed at any time: cell phones, Blackberries, iPods, PDAs, or any other electronic devices. Calculators on other devices are strictly prohibited during the exams. Information exchanges on electronic devices during class and exams are also prohibited and violate the Academic Integrity Code of Michigan Tech.
University Policies

Academic regulations and procedures are governed by University policy. Academic dishonesty cases will be handled in accordance the University's policies.

If you have a disability that could affect your performance in this class or that requires an accommodation under the Americans with Disabilities Act, please see me as soon as possible so that we can make appropriate arrangements. The Affirmative Action Office has asked that you be made aware of the following:

Michigan Tech complies with all federal and state laws and regulations regarding discrimination, including the Americans with Disabilities Act of 1990. If you have a disability and need a reasonable accommodation for equal access to education or services at Michigan Tech, please call the Dean of Students Office, at 487-2212. For other concerns about discrimination, you may contact your advisor, department head or the Affirmative Action Office, at 487-3310

Academic Integrity:
http://www.studentaffairs.mtu.edu/dean/judicial/policies/academic_integrity.html

Affirmative Action:
http://www.admin.mtu.edu/aaq/ 

Disability Services:
http://www.admin.mtu.edu/urel/studenthandbook/student_services.html#disability

Equal Opportunity Statement:
**Tentative Course Schedule**

<table>
<thead>
<tr>
<th>Main Topics</th>
<th>Chapter</th>
<th>Week</th>
</tr>
</thead>
<tbody>
<tr>
<td>Basic Concepts, 0th Law, Equation of State, Mathematical Techniques, Gas Laws, Ideal vs Real Gases</td>
<td>1, 8</td>
<td>1-2</td>
</tr>
<tr>
<td>Carnot Cycle, Entropy, 2nd Law of Thermodynamics</td>
<td>3</td>
<td>4</td>
</tr>
<tr>
<td>Spontaneity Conditions, Gibbs and Helmholtz Energies, Natural Variable Equations, Maxwell Relationships, Chemical Potential</td>
<td>4</td>
<td>5</td>
</tr>
<tr>
<td>Standard Thermodynamic Functions of Reactions, Hess’s Law, Kirchoff’s Law, 3rd Law of Thermodynamics</td>
<td>5</td>
<td>6</td>
</tr>
<tr>
<td>Reaction Equilibrium and Ideal Gases, Equilibrium Constant and Temperature Dependence (van’t Hoff Equation)</td>
<td>6</td>
<td>7</td>
</tr>
<tr>
<td>One-Component Phase Equilibrium: Phase Rule, Phase Diagrams, Clapeyron Equation</td>
<td>7</td>
<td>8</td>
</tr>
<tr>
<td>Solutions: Ideal Solutions and Raoult’s Law, Ideally Dilute Solutions and Henry’s Law</td>
<td>9</td>
<td>9</td>
</tr>
<tr>
<td>Non-Ideal Solutions: Activities, Excess Functions, Gas mixtures and Fugacity</td>
<td>10</td>
<td>10</td>
</tr>
<tr>
<td>Two-Component Liquid-Vapor Equilibrium</td>
<td>12</td>
<td>11</td>
</tr>
<tr>
<td>Kinetics and Rate Laws, Arrhenius, Reversible, Parallel, and Consecutive Reactions</td>
<td>16</td>
<td>13,14</td>
</tr>
</tbody>
</table>

Throughout the Course I will provide specific information about where to find, on the textbook, the topics covered in class.

Sometimes I will provide external notes that will be posted on Canvas.

Topics listed above and not covered in class will not be part of any assignment.

**More Information about Online Quizzes, Homework, Midterm and Final Exams**

Please, note that the adopted textbook contains problems at the end of each chapter that you may consider to solve even though they will be not assigned as homework. Homework and Exams may be partially based on them and/or related to them.

**Online Quizzes** will be assigned through Canvas: please, read carefully the description provided before in this document. Note that you are allowed to use books, notes, and calculators. **You are also allowed to work in groups but remember that you will be alone during Exams.** Online quizzes are specifically thought and designed to get you acquainted with the procedure that will be adopted during Midterm and Final Exams.


**Homework** will be assigned and returned in class according to the following schedule:

<table>
<thead>
<tr>
<th>Assignment</th>
<th>Assigned</th>
<th>Due Date</th>
</tr>
</thead>
<tbody>
<tr>
<td>Homework #1</td>
<td>Friday, 21&lt;sup&gt;st&lt;/sup&gt; September 2012</td>
<td>Friday, 28&lt;sup&gt;th&lt;/sup&gt; September 2012</td>
</tr>
<tr>
<td>Homework #2</td>
<td>Friday, 12&lt;sup&gt;th&lt;/sup&gt; October 2012</td>
<td>Friday, 19&lt;sup&gt;th&lt;/sup&gt; October 2012</td>
</tr>
<tr>
<td>Homework #3</td>
<td>Wednesday, 7&lt;sup&gt;th&lt;/sup&gt; November 2012</td>
<td>Wednesday, 14&lt;sup&gt;th&lt;/sup&gt; November 2012</td>
</tr>
<tr>
<td>Homework #4</td>
<td>Wednesday, 28&lt;sup&gt;th&lt;/sup&gt; November 2012</td>
<td>Wednesday, 5&lt;sup&gt;th&lt;/sup&gt; December 2012</td>
</tr>
</tbody>
</table>

Homework can be solved with the support of books, notes, calculators. Students are allowed to work in groups.

**Midterm Exams** are scheduled for:

1) **Wednesday, 3<sup>rd</sup> October 2012, DOW 641, 1:05 pm - 1:55 pm**

2) **Wednesday, 31<sup>st</sup> October 2012, DOW 641, 1:05 pm - 1:55 pm**

3) **Friday, 30<sup>th</sup> November 2012, DOW 641, 1:05 pm - 1:55 pm**

Midterm Exams can be solved with the support of calculators only. No books or notes will be allowed. I will provide the formula sheet.

**Final Exam** is not scheduled at the time this Syllabus has been completed. The final exam will be scheduled by the Registrar’s office later on in the semester.

The Final Exam can be solved with the support of calculators only. No books or notes will be allowed. I will provide the formula sheet.

The Final Exam will be a comprehensive 2-hour examination.

The last two/three lectures of the Course, before the Final will be devoted to a general Review of the material covered.
Getting Help in Learning

From me, as your Instructor

Please, know that I am always available to my students! Therefore, do not hesitate in contacting me for any problem or for guidance regarding the material covered in class.

In particular, my office hours are scheduled after each lecture (MWF, 2:30 pm – 3:30 pm) in my office in ChemSci 701A. I am also always available to provide assistance by appointment. In this case, you can priory contact me by phone (487-1602) or by email (lvalenza@mtu.edu).

In the latter case, please type in the subject CH3510, so that I can classify your message at a higher level of priority and reply to you as soon as possible.

According to the pace we’ll be able to keep, I may arrange for Review Lectures before the three Midterms, either during the regular time or in the evening (extra-time).

Exam Review Lectures will be held for sure during the last week of class as preparation for the Final Exam.

I will constantly remind you about the provided coaching sessions (see below).

From Eric Kamischke (Math Coach)

During weeks 2, 3, and 4, Eric will help you in reviewing calculus techniques needed in Physical Chemistry for a total of four hours. In the following table dates and location are reported for these meetings:

<table>
<thead>
<tr>
<th>Location</th>
<th>Day</th>
<th>Time</th>
</tr>
</thead>
<tbody>
<tr>
<td>ChemSci 101</td>
<td>Wed, 12th September</td>
<td>7 pm – 8 pm</td>
</tr>
<tr>
<td>ChemSci 101</td>
<td>Wed, 19th September</td>
<td>7 pm – 8 pm</td>
</tr>
<tr>
<td>ChemSci 101</td>
<td>Mon, 24th September</td>
<td>7 pm – 8 pm</td>
</tr>
<tr>
<td>ChemSci 101</td>
<td>Tue, 25th September</td>
<td>7 pm – 8 pm</td>
</tr>
</tbody>
</table>

From Morgan Parr, through the Chemistry Learning Center (CLC)

Throughout the entire Course, Morgan will provide you help with tutorial sessions. Note that these meetings are not supposed to be solution giving sessions, but stepping stone sessions to give guidance in the right direction for solving problems and improving your reasoning skills.

The P-Chem sessions are not compulsory. Nevertheless, you are warmly invited to attend them as part of your learning process. They will provide good help in acquiring the knowledge and capabilities needed to master the material covered in class.

In addition, the P-Chem sessions will be useful in getting some guidance in solving the assigned homework and in reviewing the material before Midterm and Final Exams.
Once again, note that the adopted textbook contains problems at the end of each chapter that you may consider to solve even though they will be not assigned as homework. Homework and Exams may be partially based on them and/or related to them.

Sessions will be held in the Chemistry Learning Center (CLC), located on the 2\textsuperscript{nd} floor of the Chemical Sciences Building in Room 208. They will run every Tuesday and Thursday, 7:00 pm – 9:00 pm, starting from the week 2 until week 14. The Chemistry Learning Center (CLC) is a free service provided by the Department of Chemistry and the University thought as a valuable resource for students enrolled in PChem. The CLC is staffed by upper level undergraduates (coaches) who have a good background in chemistry and are familiar with the course.
Excused Absences

Events beyond your control may cause you to miss a homework deadline or an exam. Whenever possible, contact me prior to your absence to arrange to make-up missed work. If you are unable to notify me concerning an absence or if you need to notify several instructors on short notice, contact the Office of Student Affairs for assistance. The Dean of Students will then inform all your instructors that you face a situation that requires that you miss class, and you are granted an excused absence. It is then your responsibility to contact each of your instructors after you recover from your illness or return to campus.

An absence is excused under the following conditions:

- If you participate in off-campus University-sponsored activities such as field trips, fine arts performances, intercollegiate athletics, job fairs, etc., you are granted an excused absence if your activity conflicts with an exam. Furthermore, I consider plant trips, job interviews requiring travel, and professional society meetings as excusable. It is imperative that for an absence of this type, for which a conflict with an exam is known well ahead of time, that you arrange with me to take the exam earlier than its normally scheduled time.

- If you encounter circumstances beyond your control such as illness, the funeral of any relative or close friend, or other personal emergency, you are granted an excused absence. You must provide verification of the special circumstances that led to your absence. In the event of a missed exam due to an excused absence, it is not possible to make-up the exam. Instead, an excused absence from an exam will receive the score EX. At the end of the semester, exam EX scores will be replaced by a weighted average of all of your non-EX scores on exams (midterms and final exams). If the final exam is missed as a result of an excused absence, you will be awarded the letter grade of I (incomplete) and must take the CH 3510 final exam at the end of any one of the next semesters that you’re in residence. Two or more exams missed as a result of excused absences will be handled on an individual basis.

If a homework due date is missed as a result of an excused absence, the due date will be extended after you notify me.